#### Nuclear Magnetic Resonance Spectroscopy

#### Features:

- Used to identify products of reactions
- Also gives information about chemical environment, connectivity and bonding of nuclei

#### **Requirements:**

- Pure or mostly pure sample of material (not useful for analyzing mixtures)
- > 1 mg material

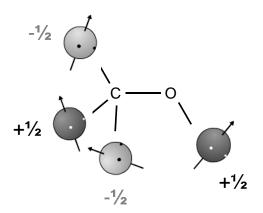
#### Nuclear Magnetic Resonance Spectroscopy Depends on Nuclear Spin

| Nucleus         | <u>Spins</u>      | Isotope<br>Abundance |               | spin |
|-----------------|-------------------|----------------------|---------------|------|
| <sup>1</sup> H  | <del></del> _     | 99.9%                |               | +1/2 |
|                 | $\pm \frac{1}{2}$ |                      |               | 1 /2 |
| <sup>12</sup> C | 0                 | 98.9%                | \             |      |
| <sup>13</sup> C | ±1/2              | 1.1%                 |               |      |
| <sup>19</sup> F | ±1/2              | 100%                 |               | -1/2 |
| <sup>31</sup> P | ±½                | 100%                 | $\mathcal{I}$ |      |

We'll be talking about spectroscopy of <sup>1</sup>H (and later <sup>13</sup>C).

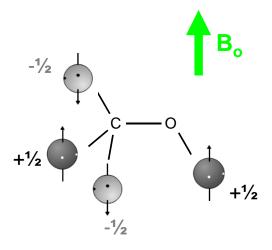
#### Each <sup>1</sup>H Nucleus Has Its Own Spin

We'll consider <sup>1</sup>H nuclei in CH<sub>3</sub>OH as an example.



In the absence of an applied field, nuclei are randomly oriented.

# Nuclear Spins Are Aligned By An Applied Magnetic Field

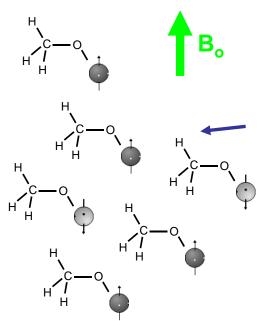


In the absence of an applied field, nuclei are randomly oriented.

In the presence of an applied field  $\mathbf{B}_{\mathbf{o}}$ , nuclei are oriented by the field.

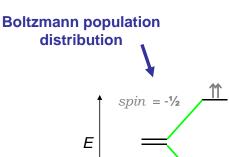
- + spins align with the field,
- spins align against the field.

# Overall, Applied Field Leads to Bulk Spin Magnetization

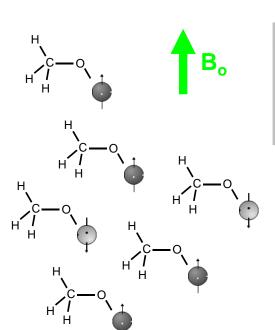


Being aligned with field is more stable than against field.

So, more spins align with the field.



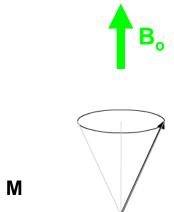
# Overall, Applied Field Leads to Bulk Spin Magnetization



Being aligned with field is more stable than against field.

So, more spins align with the field, and sum of spins also aligns with the field.

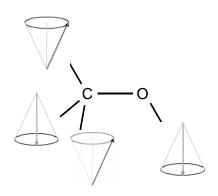
#### Spins, And Overall Spin, Precess (Wobble) in Applied Magnetic Field



bulk

magnetization

Again, using CH<sub>3</sub>OH as an example,



All nuclei wobble at a characteristic frequency.

#### **Spins, And Overall Spin, Precess** (Wobble) in Applied Magnetic Field



Nuclei precess at the Larmor frequency.  $v_o = \frac{\gamma B_o}{2\pi}$ 

$$v_o = \frac{\gamma B_o}{2\pi}$$

M

bulk magnetization B 4.73 T 9.46 T 11.75 T  $v_o$  (<sup>1</sup>H) 200 MHz 400 MHz 500 MHz  $v_0$  (13C) 50 MHz 75 MHz 125 MHz

(Precession is <u>fast</u>.)

### Big Fields Means Big Magnets



Interior of a 4.73 T magnet (on display in NMR Facility, Smith Hall)

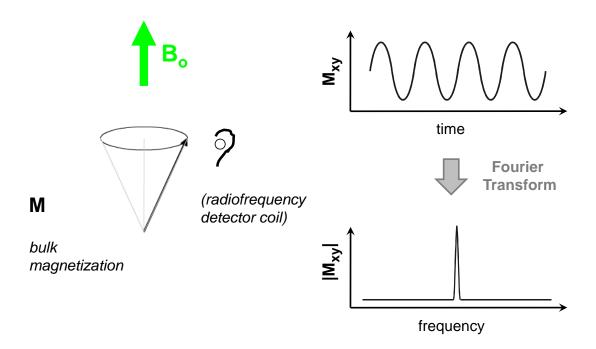


Installation of 16.45 T magnet in Hasselmo Hall.

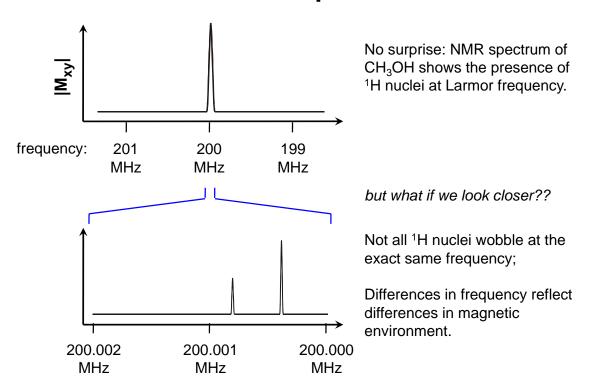
### **An NMR Facility**



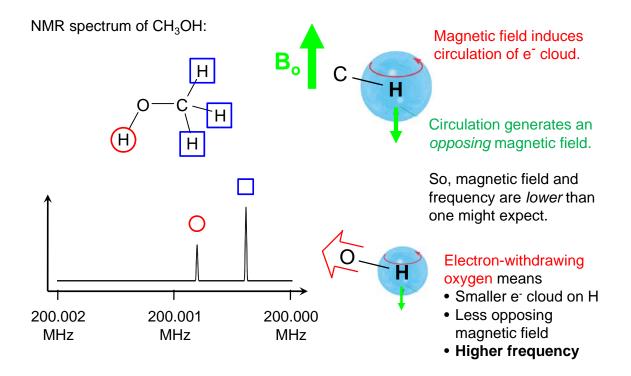
# An NMR Spectrometer "Listens" to Frequency of Nuclear Precession



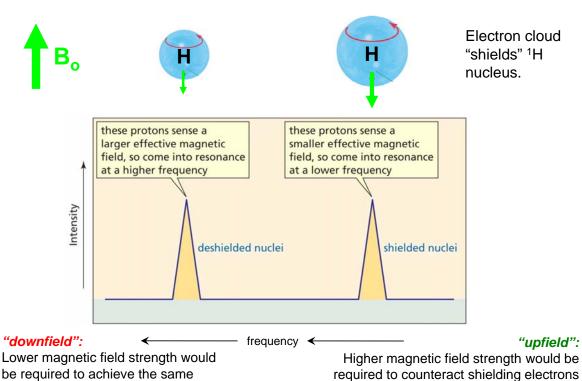
#### **An NMR Spectrum**



#### "Shielding" Influences Proton Frequency



#### "Shielding" Influences Proton Frequency



and achieve the same frequency.

frequency.

# **Chemical Shift:**A Proportional Horizontal Axis

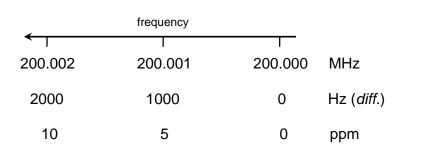
**Problem:** Differences in frequency depend on spectrometer field

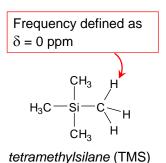
strength, vary from instrument to instrument.

**Solution:** Define an absolute scale independent of spectrometer

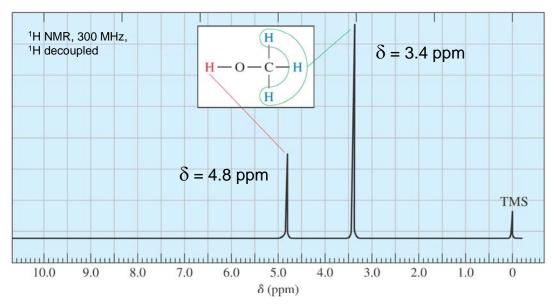
frequency, called "chemical shift".

chemical shift, ppm  $\delta = \frac{\text{shift downfield from TMS (in Hz)}}{\text{spectrometer frequency (in MHz)}}$ 





## Chemical Shift: A Proportional Horizontal Axis



CH<sub>3</sub>OH protons have these chemical shifts (ppm values), regardless of instrument they are measured on.

### **Different Types of Protons Have Characteristic Chemical Shifts**

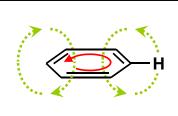
TABLE 13-3 Typical Values of Chemical Shifts

| Type of Proton              | Approximate $\delta$ | Type of Proton     | Approximate $\delta$  |
|-----------------------------|----------------------|--------------------|-----------------------|
| alkane (—CH <sub>3</sub> )  | 0.9                  | C=CCCH3            | 1.7                   |
| alkane (—CH <sub>2</sub> —) | 1.3                  | CH <sub>3</sub>    |                       |
| alkane (—CH—)               | 1.4                  | Ph—H               | 7.2                   |
| alkane (—CH—)               |                      | Ph—CH <sub>3</sub> | 2.3                   |
| O                           |                      | R—CHO              | 9-10                  |
| -C-CH <sub>3</sub>          | 2.1                  | R-COOH             | 10-12                 |
| $-C \equiv C - H$           | 2.5                  | R—OH               | variable, about 2-5   |
| $R-CH_2-X$                  | 3-4                  | Ar—OH              | variable, about 4-7   |
| (X = halogen, O)            |                      | R—NH,              | variable, about 1.5-4 |
| c=c                         | 5-6                  | 2                  | , minore, about 110   |

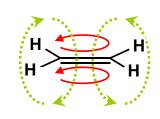
A better resource: http://www.chem.wisc.edu/areas/reich/Handouts/nmr-h/hdata.htm

### Chemical Shift: Multiple Bonds and Induced Current

In presence of applied field

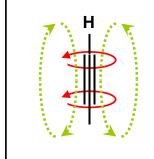


"Ring current" reinforces applied  $B_0$  field; **deshields** aromatic protons  $(\delta = 6.5-8.5 \text{ ppm}).$ 



Induced current is less effective in olefins, still **deshields**, but less  $(\delta = 4-6 \text{ ppm})$ .





Acetylene protons are **shielded**  $(\delta = 1.5-3 \text{ ppm}).$