

NAME _____

ID # _____

ORGANIC CHEMISTRY II (2302)

8:00 – 8:50 am, July 23, 2015

Exam 3

If you want to pick this exam up Monday in class (in public), please check the box on the right:

If you do not check the box, I will not bring your exam to class on Monday, and you will need to pick up your exam in private from Chemistry department staff in 115 Smith beginning Tuesday, July 28th. Exams that are not picked up within two weeks will be disposed of.

A periodic table and a chart of reaction conditions are attached to the back of this exam as an aid. Otherwise, you are not permitted to use any other materials (including notes, books, or electronic devices of any kind).

Right now, write your name and student ID number at the top of this page. When the exam begins, please write your name at the top of the next page.

You may use pen or pencil. However, re-grades will be considered only for exams completed in pen.

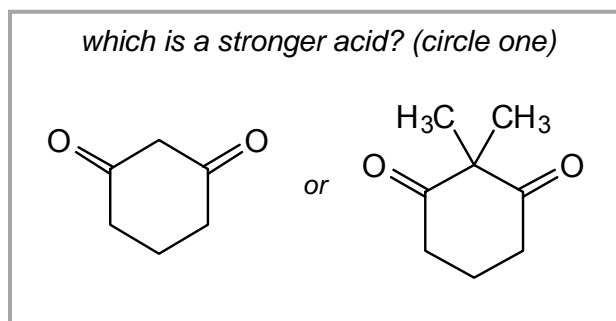
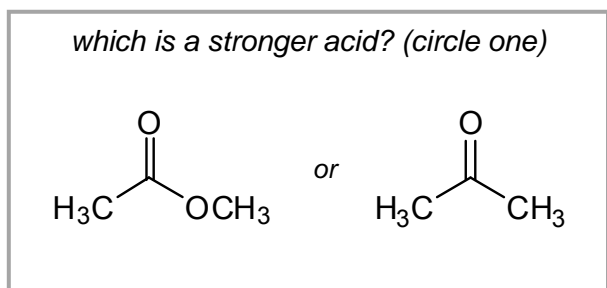
Please write your answers in the boxes/spaces provided. If your answer is not in the appropriate space (say, for example, it's on the back of the page), draw us an arrow and/or note telling us where to look.

NAME _____

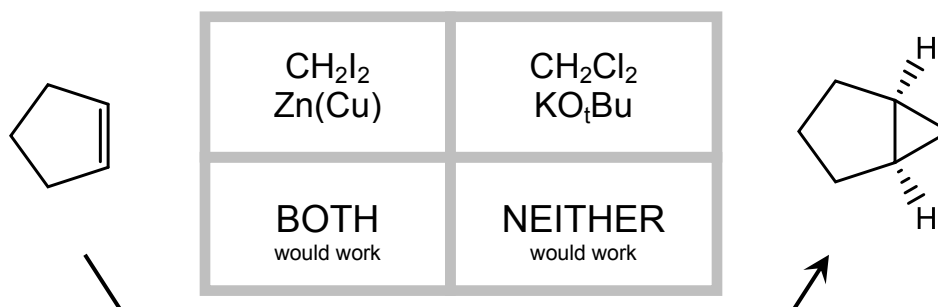
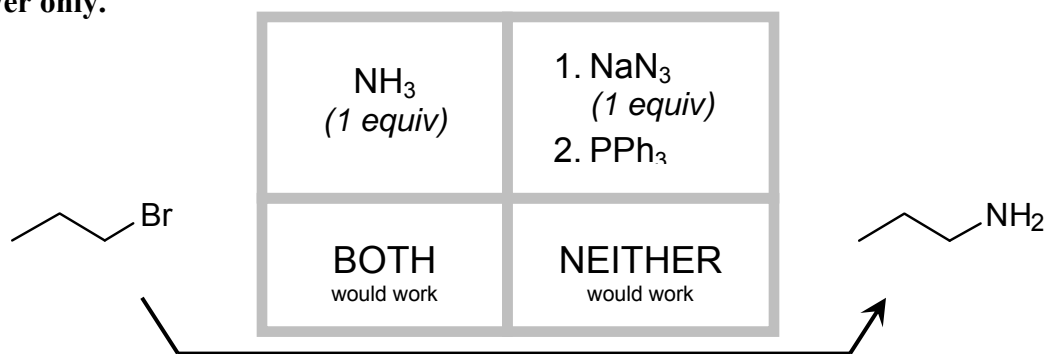
Scoring: 1. _____ / 6 4. _____ / 24
 2. _____ / 20 5. _____ / 34
 3. _____ / 16

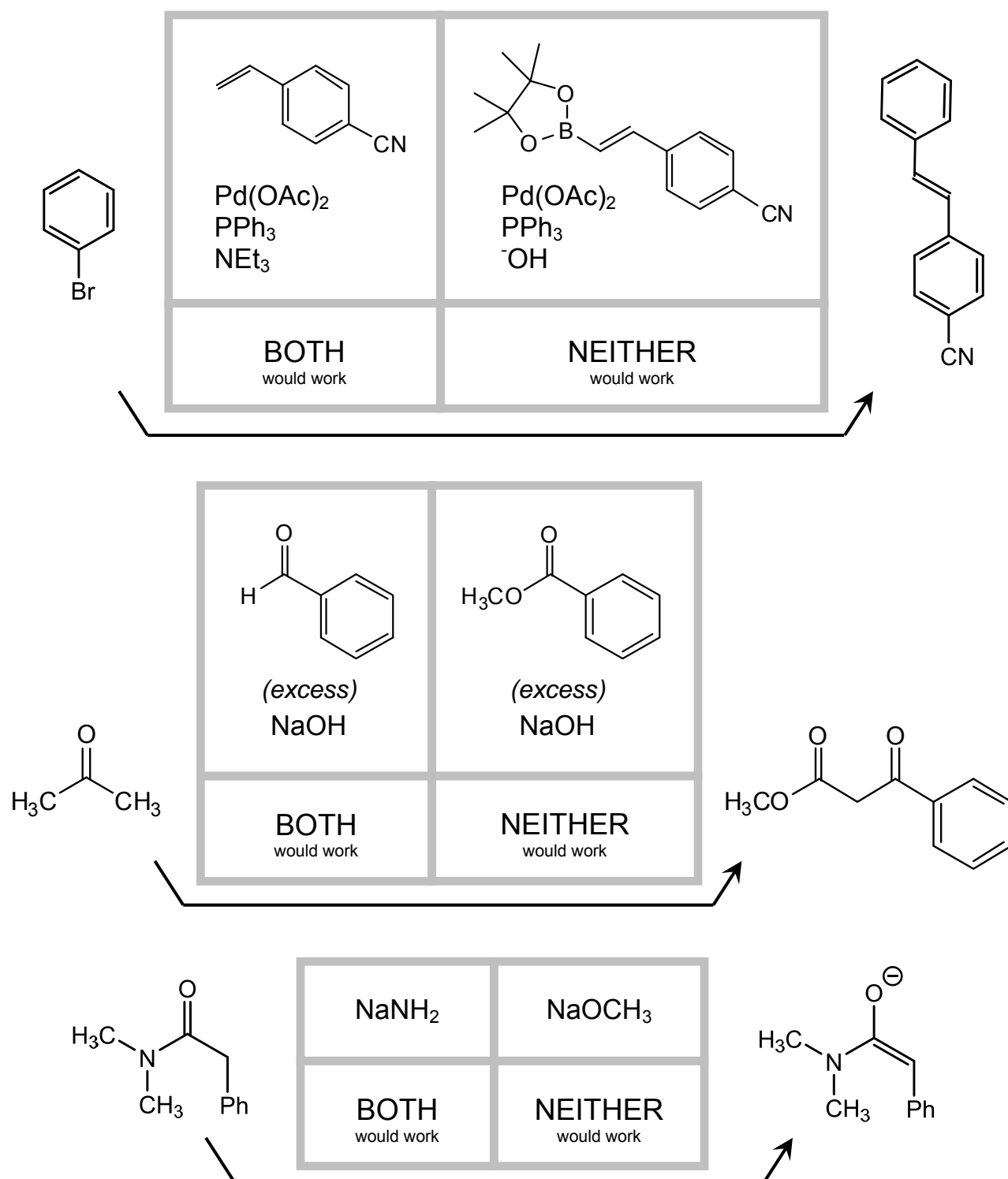
Total Score: _____ / 100

1. (6 pts) Each of the carbonyl-containing compounds below is acidic, and can be deprotonated to form an enolate anion. For each pair, **circle the stronger acid**.

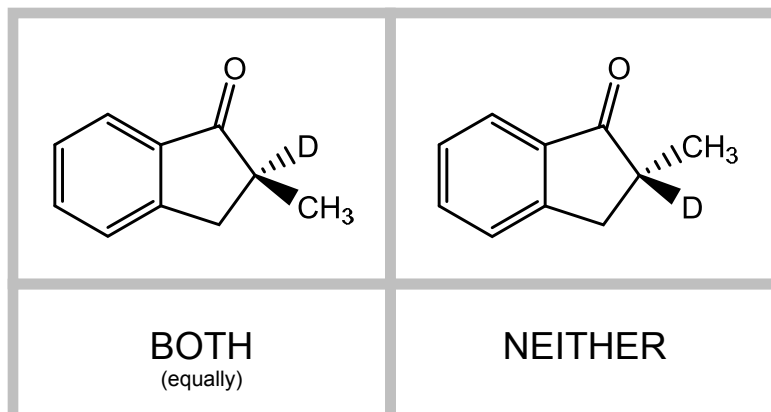
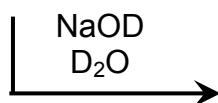
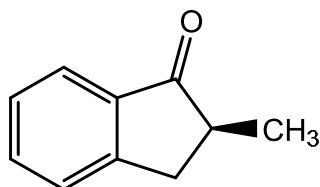
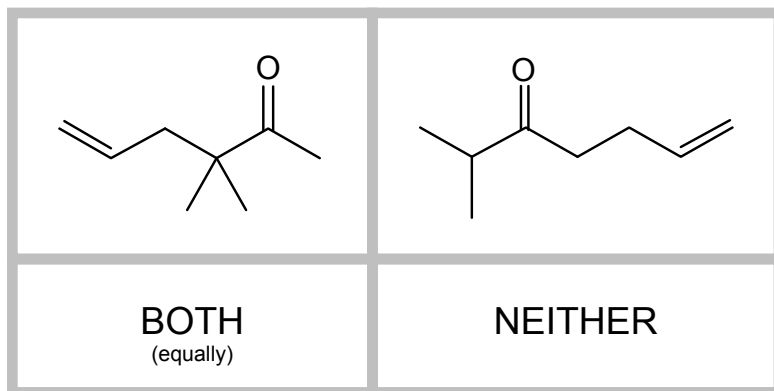
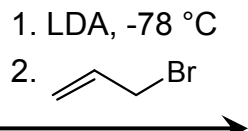
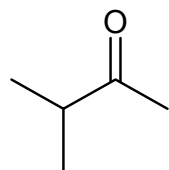
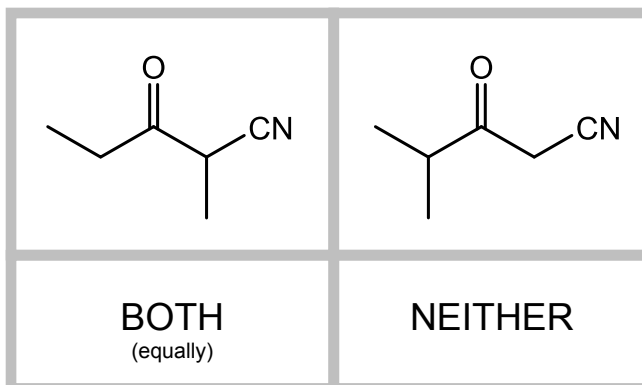
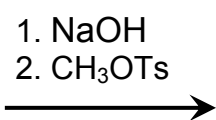
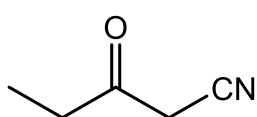
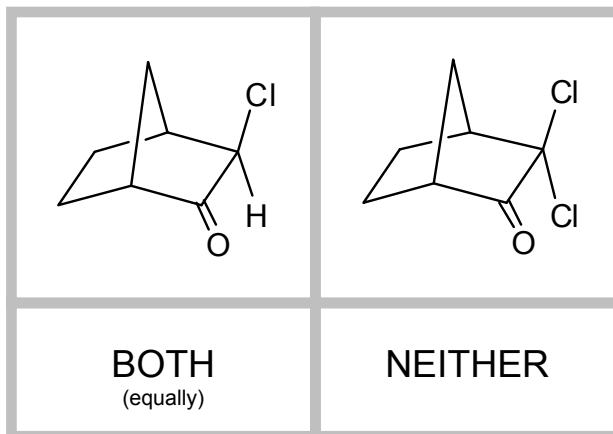
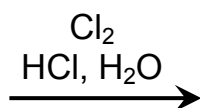
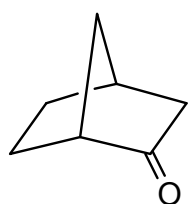


2. (20 pts) Each of the reactions below is drawn with two possible reaction conditions. If only one of the two reaction conditions would generate the given molecule as the major product, circle those conditions. If both sets of conditions would accomplish the reaction, circle "BOTH". If neither set of reaction conditions would succeed, circle "NEITHER". **Circle one answer only.**





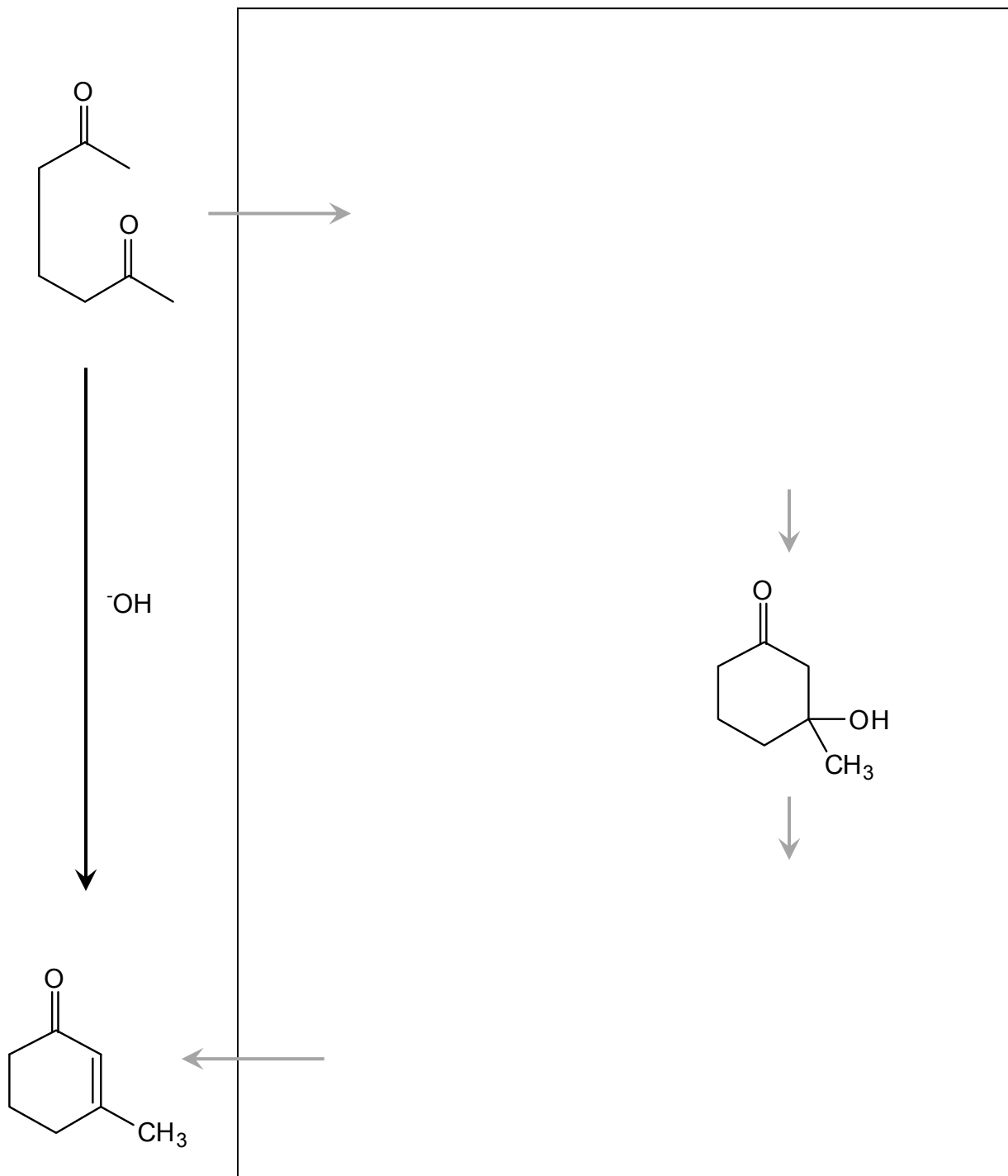
3. (16 pts) Each of the reactions on the next page is drawn with two possible products. If one of the two products predominates, circle that preferred product. If the two products are produced equally, circle "BOTH". If neither product would result from the reaction, circle "NEITHER". **Circle one answer only.**



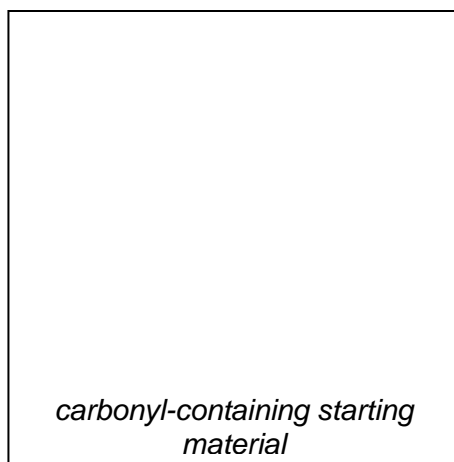
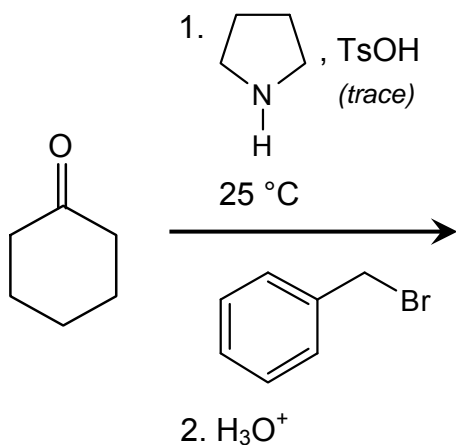
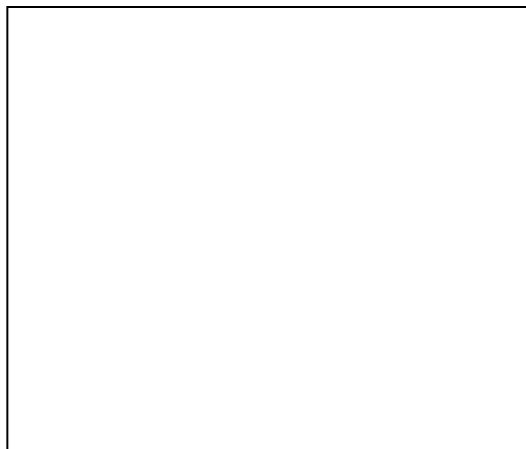
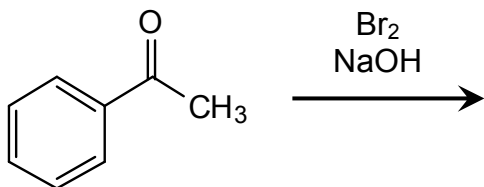
4. (24 pts) For the reaction shown below, draw a mechanism that explains how the product is generated from the starting material. In your answer, make sure that you:

- Draw each step of the mechanism separately;
- Use “electron pushing” to show where the electrons in each step go;
- Use only the molecules that you are given; do not invoke reactants or solvents that aren't in the problem.

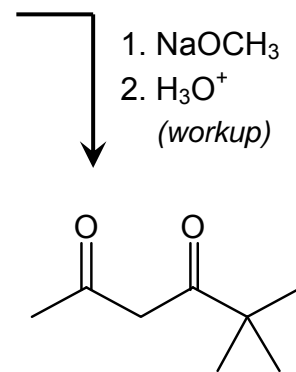
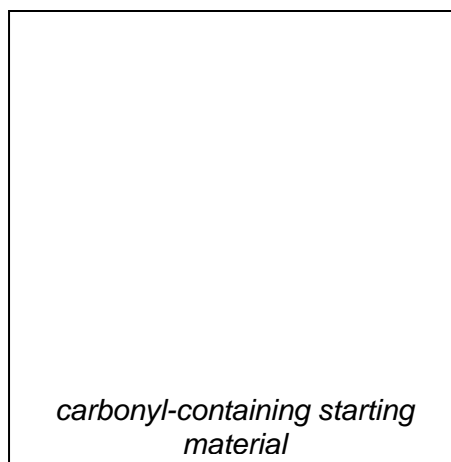
I have drawn one intermediate in the mechanism for you, but you will still need to push electrons for steps to and from that intermediate.



5. (34 pts) For each of the reactions on the following pages, fill in the empty box corresponding to reactants, reagents, or products. Give only one answer in each box. For reactions that you expect to yield multiple products, draw one major product. For reactions that yield multiple enantiomers, draw only one enantiomer in the box, and include the note "+ enantiomer".



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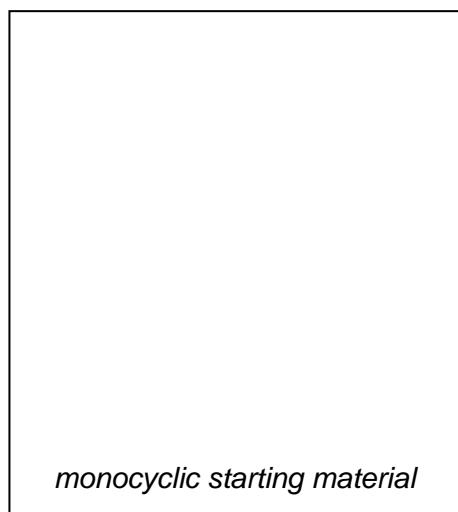
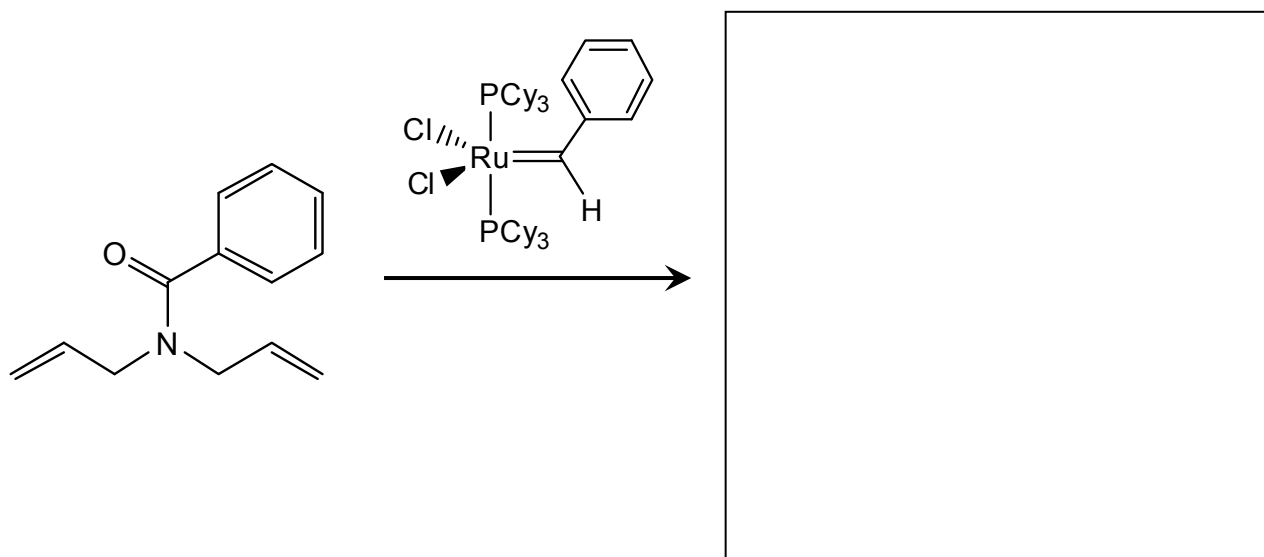
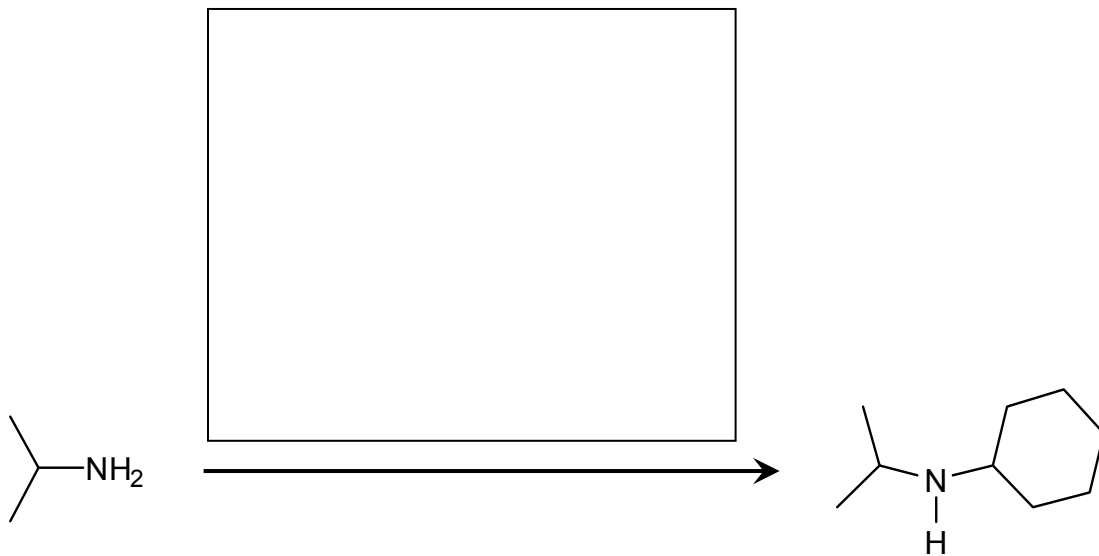
used in excess?

used in excess?

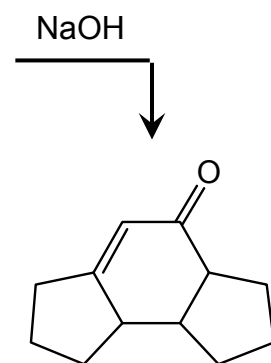
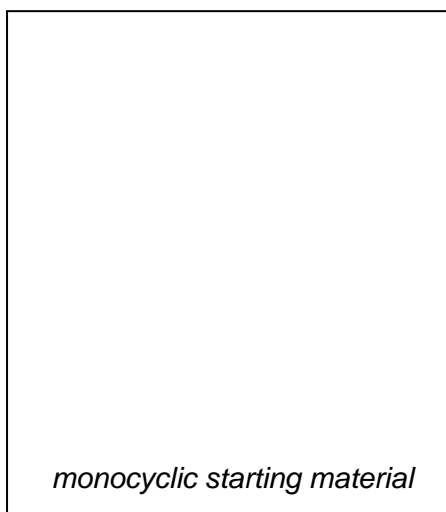
YES or NO ?

YES or NO ?

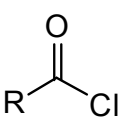
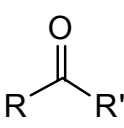
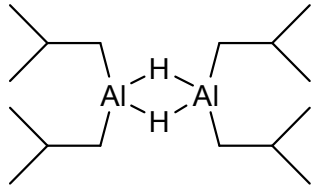
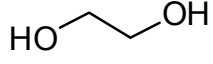
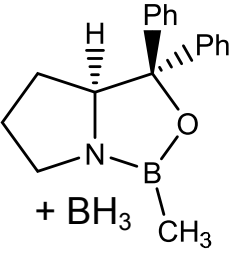
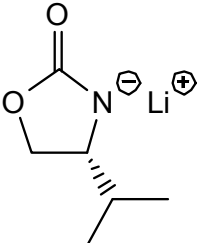
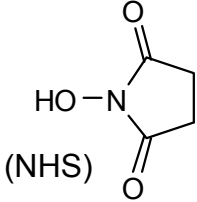
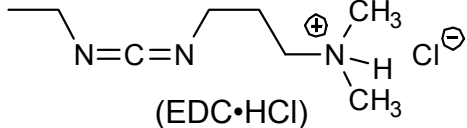
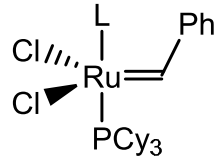
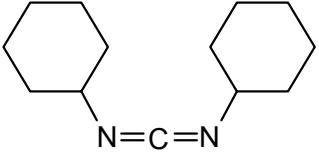
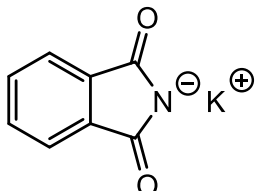
(circle one)



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Exam 3 Chart of Reaction Conditions

Br ₂ FeBr ₃	Cl ₂ AlCl ₃	H ₂ SO ₄ HNO ₃	Sn or Fe HCl/H ₂ O	H ₂ SO ₄ SO ₃	KMnO ₄ OH ⁻ , 100 °C
1. NaNO ₂ HCl 2. CuCN or H ₃ PO ₂ or CuX or H ₃ O ⁺	 AlCl ₃	R-X (R = alkyl) AlCl ₃ or FeBr ₃	Zn(Hg), HCl/H ₂ O	1. N ₂ H ₄ 2. KOH, Δ	
	RMgX	H ₂ Pd-C	Li hexane	Mg Et ₂ O	R ₂ CuLi RLi
Na ₂ Cr ₂ O ₇ H ₂ SO ₄	1. NaBH ₄ 2. H ₂ O	LiAlH(OtBu) ₃	1. PPh ₃ 2. n-BuLi 3. 	 (DIBAL-H)	
1. O ₃ 2. H ₂ O	1. LiAlH ₄ 2. H ₂ O	Bu ₄ N ⁺ F ⁻	1. Ag ₂ O NH ₃ 2. H ₃ O ⁺		
 HCl	NaNH ₂	(COCl) ₂	PhCH ₂ Br Ag ₂ O	 + BH ₃ CH ₃	
	SOCl ₂ (& pyridine, usually)	(CH ₃) ₃ SiCl {TMSCl}, or TBDMSCl; Et ₃ N or imidazole			
1.  2. R-X 3. LiOH	 (NHS)	 (EDC·HCl)	1. CH ₃ I (excess) 2. Ag ₂ O H ₂ O		
 (Grubbs catalyst)	 (DCC)	1.  2. N ₂ H ₄ (or OH ⁻)	RCHO Na(OAc) ₃ BH or NaBH ₃ CN		
	CHCl ₃ KOtBu	1. NaN ₃ 2. PPh ₃ H ₂ O	Pd(PPh ₃) ₄ NaOH	Pd(OAc) ₂ PPh ₃ , NEt ₃	
			CH ₂ I ₂ Zn(Cu)		

		1		2		3		4		5		6		7		8		9		10		11		12		13		14		15		16		17		18																																																																																																																																																																																																				
		1A		2A		3B		4B		5B		6B		7B		8B						1B		2B		3A		4A		5A		6A		7A		8A																																																																																																																																																																																																				
1	1	H Hydrogen 1.01	2	He Helium 4.00	3	4	Li Lithium 6.94	5	Be Beryllium 9.01	6	7	B Boron 10.81	8	C Carbon 12.01	9	N Nitrogen 14.01	10	O Oxygen 16.00	11	F Fluorine 19.00	12	Ne Neon 20.18	13	Na Sodium 22.99	14	Mg Magnesium 24.31	15	Al Aluminum 26.98	16	Si Silicon 28.09	17	P Phosphorus 30.97	18	S Sulfur 32.07	19	Cl Chlorine 35.45	20	Ar Argon 39.95	21	K Potassium 39.10	22	Ca Calcium 40.08	23	Sc Scandium 44.96	24	Ti Titanium 47.87	25	V Vanadium 50.94	26	Cr Chromium 52.00	27	Mn Manganese 54.94	28	Fe Iron 55.85	29	Co Cobalt 58.93	30	Ni Nickel 58.69	31	Cu Copper 63.55	32	Zn Zinc 65.39	33	Ga Gallium 69.72	34	Ge Germanium 72.61	35	As Arsenic 74.92	36	Se Selenium 78.96	37	Rb Rubidium 85.47	38	Sr Strontium 87.62	39	Y Yttrium 88.91	40	Zr Zirconium 91.22	41	Nb Niobium 92.91	42	Mo Molybdenum 95.94	43	Tc Technetium (98)	44	Ru Ruthenium 101.07	45	Rh Rhodium 102.91	46	Pd Palladium 106.42	47	Ag Silver 107.87	48	Cd Cadmium 112.41	49	In Indium 114.82	50	Sn Tin 118.71	51	Sb Antimony 121.76	52	Te Tellurium 127.60	53	I Iodine 126.90	54	Xe Xenon 131.29	55	Cs Cesium 132.91	56	Ba Barium 137.33	57	La Lanthanum 138.91	58	Ce Cerium 140.12	59	Pr Praseodymium 140.91	60	Nd Neodymium 144.24	61	Pm Promethium (145)	62	Sm Samarium 150.36	63	Eu Europium 151.96	64	Gd Gadolinium 157.25	65	Tb Terbium 158.93	66	Dy Dysprosium 162.50	67	Ho Holmium 164.93	68	Er Erbium 167.26	69	Tm Thulium 168.93	70	Yb Ytterbium 173.04	71	Lu Lutetium 174.97	72	Fr Francium (223)	73	Ra Radium (226)	74	Ac Actinium (227)	75	Rf Rutherfordium (261)	76	Hf Hafnium 178.49	77	Ta Tantalum 180.95	78	W Tungsten 183.84	79	Re Rhenium 186.21	80	Os Osmium 190.23	81	Ir Iridium 192.22	82	Pt Platinum 195.08	83	Au Gold 196.97	84	Hg Mercury 200.59	85	Tl Thallium 204.38	86	Pb Lead 207.2	87	Bi Bismuth 208.98	88	Po Polonium (209)	89	At Astatine (210)	90	Rn Radon (222)	91	Th Thorium 232.04	92	Pa Protactinium 231.04	93	U Uranium 238.03	94	Np Neptunium (237)	95	Pu Plutonium (244)	96	Am Americium (243)	97	Cm Curium (247)	98	Bk Berkelium (247)	99	Cf Californium (251)	100	Fm Fermium (257)	101	Md Mendelevium (258)	102	No Nobelium (259)	103	Lr Lawrencium (262)	104	Rf Rutherfordium (261)	105	Sg Seaborgium (266)	106	Bh Bohrium (264)	107	Hs Hassium (269)	108	Mt Meitnerium (268)	109	Ds Darmstadtium (271)	110	Rg Roentgenium (272)	111	Cn Copernicium (285)	112	Nh Nihonium (284)	113	Fl Flerovium (287)	114	Mc Moscovium (288)	115	Lv Livermorium (293)	116	Ts Tennessine (294)	117	Og Oganesson (294)

Key

11	Na	Sodium	22.99
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— Atomic number
— Element symbol
— Element name

Average atomic mass*

* If this number is in parentheses, then it refers to the atomic mass of the most stable isotope.