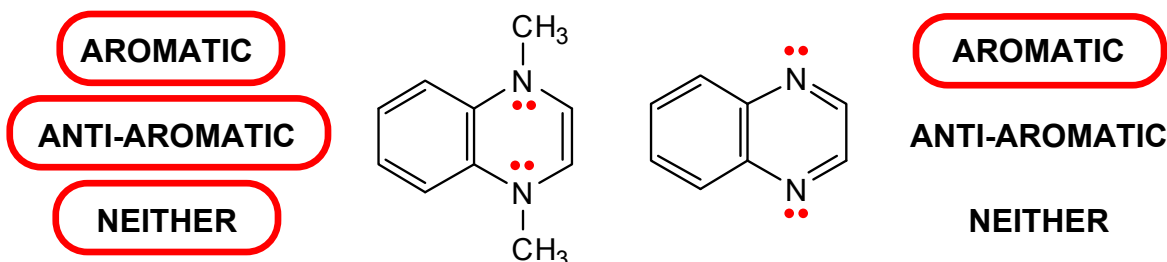
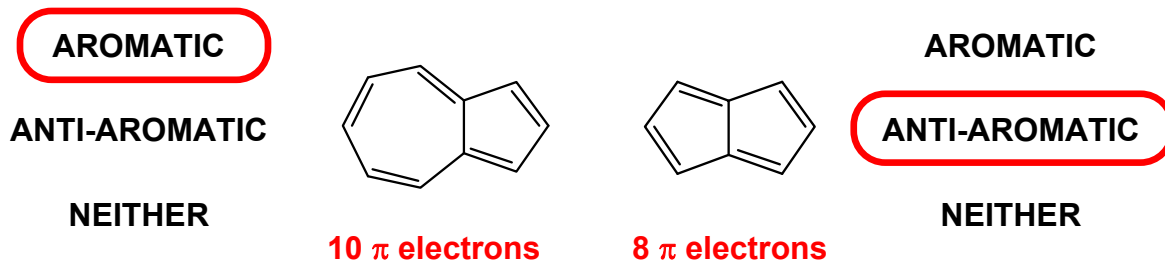


Workshop 4 Solutions
Aromatic, Or Not?

1. For each of the following molecules, circle whether the molecule is aromatic, anti-aromatic, or neither.

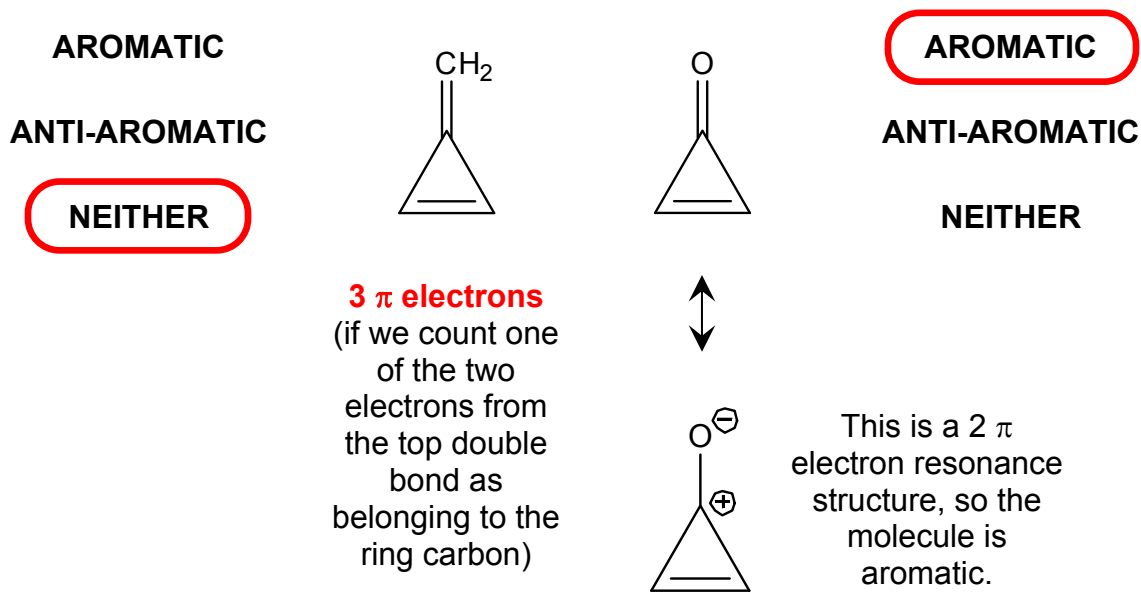
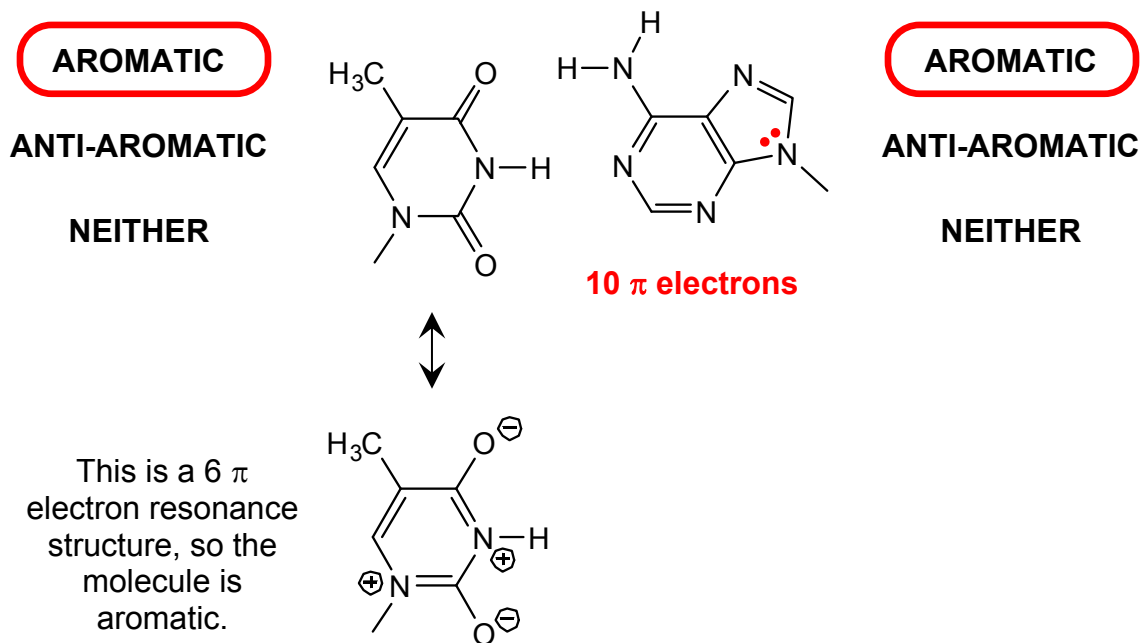


If the molecule were flat, with sp^2 -hybridized nitrogens, then it would have **12 π electrons** and be **anti-aromatic**.

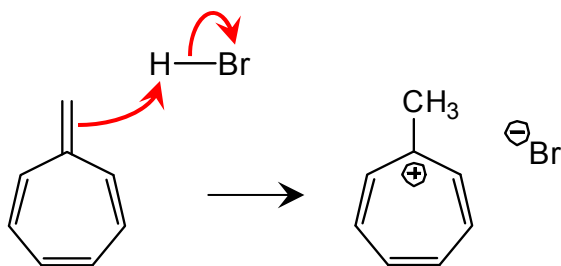
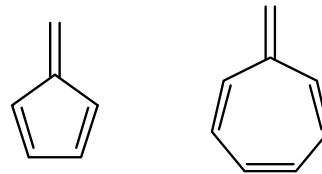
But the nitrogens can stay sp^3 -hybridized if they choose. (Seems likely they would.) if so, the right hand ring would be **non-aromatic**, and the left-hand (benzene) ring would be **aromatic**.

sp^2 ,

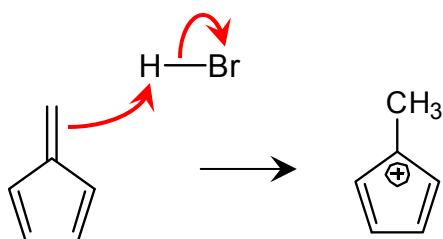
10 π electrons
(both nitrogens are sp^2 with only two σ bonds, so lone pairs are in third sp^2)



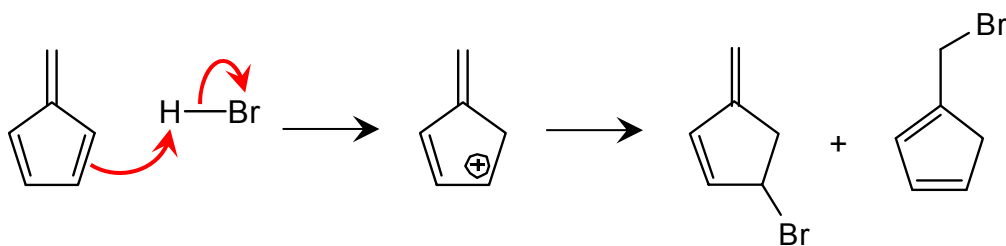
2. One of the molecules on the right will react with aqueous HBr just one way, to yield one product, selectively. The other will react with one molecule of aqueous HBr in a variety of ways, to yield a variety of products. Which molecule has which reactivity pattern, and what HBr adducts would you expect from each?



6 π electrons, aromatic. So this cation is formed specifically. In addition, the cation is stable as a salt in water, so the Br^- doesn't even add; this is the product.



4 π electrons, anti-aromatic. So this cation would not be formed. Instead, HBr would add somewhere else on the molecule, to generate a cation that would be stabilized by resonance but not anti-aromatic:



(+ others from other protonation locations)