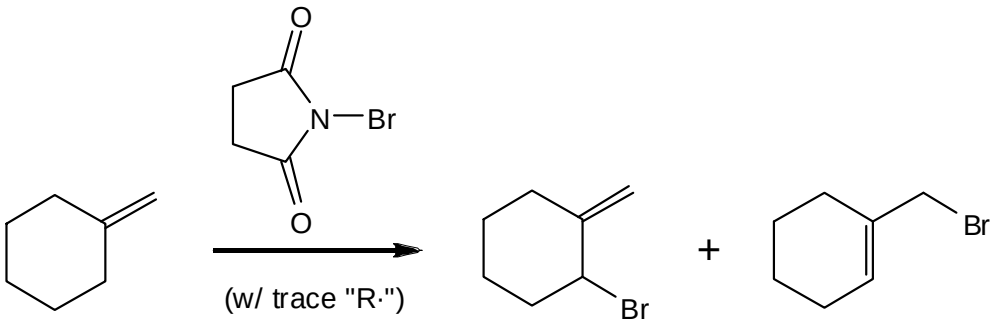
**Chemistry 2331H** Tuesday, October 29

**Workshop 17**
**Unpaired Electron Pushing**

Radical bromination of allylic C-H's is usually performed with *N*-bromosuccinimide (NBS), in order to avoid the bromine addition to the double bond that would occur if the reaction were performed with $Br_2$ alone.




a. Draw a free-radical chain reaction mechanism for this process. In the mechanism, The NBS N-Br will play the same role as Br-Br in radical bromination, and the trace "R•" means there is no need for an initiation step. (In other words, start your mechanism by using R• to take an H from the starting material.)

b. Which of the two products would you expect to predominate? Why?